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**Abstract:** Salt, is a compound, plays a role Salts are formed through a chemical reaction between an acid and a base. This reaction is called neutralization. The chemical name of a salt typically consists of the name of the cation (from the base) followed by the name of the anion (from the acid). Here's a general way to derive the chemical name of a salt: Identify the cation and anion: The cation is the positively charged ion, which comes from the base in the reaction. The anion is the negatively charged ion, which comes from the acid in the reaction. Write the name of the cation first, followed by the name of the anion. For example: Sodium ( $\text{Na}^+$ ) from sodium hydroxide ( $\text{NaOH}$ ) combines with chloride ( $\text{Cl}^-$ ) from hydrochloric acid ( $\text{HCl}$ ) to form sodium chloride ( $\text{NaCl}$ ). Make sure to balance the charges. For example, calcium ( $\text{Ca}^{2+}$ ) from calcium hydroxide ( $\text{Ca}(\text{OH})_2$ ) combines with carbonate ( $\text{CO}_3^{2-}$ ) from carbonic acid ( $\text{H}_2\text{CO}_3$ ) to form calcium carbonate ( $\text{CaCO}_3$ ). in various aspects of human life. Its chemical names, sodium chloride and  $\text{NaCl}$ , reflect its fundamental composition. This crystalline substance, composed of the sodium cations and chloride anions, that exhibits a wide range of properties, making it essential for culinary, industrial, and biological applications. This abstract explores the chemical as well as structural properties, and diverse uses of salt in different fields, highlights its significance in everyday life.

**Key words.** Salts, Types, Properties, Hydrolysis, Double Salts, Rochelle's Salt.

### **Introduction.**

Without having much knowledge about salt chemistry most people would be able to give a good salt definition. The salt definition in chemistry shows that it is an inorganic chemical that is abundantly available on the Earth's crust. It is an ionic substance composed of ions that are joined together with an electrostatic ionic bonds.

### **Main Body.**

Salt is ionic compound that has a cation other than  $\text{H}^+$  and an anion other than  $\text{OH}^-$  and is obtained along with water in the neutralization reaction between acids and bases.

Eg:-  $\text{NaCl}$ ,  $\text{CuCl}_2$  etc. Sodium chloride is one of the best-known salt.

### **Types of Salts:**

1. Acidic salt – The salt formed by partial neutralization of a diprotic or a polyprotic acid is known as an acidic salt. These salts have ionizable  $\text{H}^+$  ion along with another cation. Mostly the ionizable  $\text{H}^+$  is a part of the anion. Some acid salts are used in baking.

For eg:-  $\text{NaHSO}_4^-$ ,  $\text{KH}_2\text{PO}_4$  etc.

2. Basic or Alkali Salt – The salt formed by the partial or some neutralization of a strong base by a weak acid is known as a basic salt. They hydrolyze to form a basic solution. It is because when hydrolysis of basic salt takes place, the conjugate base of the weak acid is formed in the solution.

For eg:- White lead ( $2\text{PbCO}_3 \cdot \text{Pb}(\text{OH})_2$ ).

3. Double salt – The salts that contain more than one cation or anion are known as double salt. They are obtained by the combination of two different salts crystallized in the same ionic lattice.

For eg:- Potassium sodium tartrate ( $\text{KNaC}_4\text{H}_4\text{O}_6 \cdot 4\text{H}_2\text{O}$ ) also known as Rochelle salt.

4. Mixed Salts – The salt that consists of the fixed proportion of two salts, often sharing either a common cation or common anion is known as mixed salt.

For e.g. :-  $\text{CaOCl}_2$

#### Properties of Salts

The compound's sodium chloride has different properties from the elements sodium and chlorine.

Saltwater contains ions and is a fairly good conductor of electricity.

This electrostatic force of attraction holds the ions together and a chemical bond is said to form between them.

#### Hydrolysis of Salt:

Hydrolysis of salt is the reaction of salt with water. It is the reverse of a neutralization reaction. In this reaction, when salt undergoes reaction with water, the constituent acid and base are formed as products. In hydrolysis, the salt dissociates to form ions, completely or partially depending upon the solubility product of that salt.

#### Common Example

the following are the examples

Basic salts:-  $\text{Na}_2\text{CO}_3$ ,  $\text{NaHCO}_3$ ,  $\text{Ca}(\text{HCO}_3)_2$ ,  $\text{CaCO}_3$ ,

Normal Salts:-  $\text{NaCl}$ ,  $\text{KCl}$ ,  $\text{KNO}_3$ ,  $\text{MgCl}_2$ ,  $\text{NH}_4\text{Cl}$ ,

Complex Salts:-  $\text{Na}[\text{Ag}(\text{CN})_2]$ ,  $\text{K}_4[\text{Fe}(\text{CN})_6]$ ,  $\text{K}_3[\text{Fe}(\text{CN})_6]$ ,

#### Conclusion.

Salts in chemistry are ionic compounds that are formed through the cations complete transfer of electrons to the anions. They are formed due to a neutralisation reaction between acids and bases. Salt is classified into different types acidic, basic, normal, and double.

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